



برعاية معالي وزير التربية والتعليم
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وتوجيهات مساعد الوزير لشئون تطوير المناهج التعليمية
والمشرف على الادارة المركزية لتطوير المناهج
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Chemistry
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الإدارة المركزية لتطوير المناهج
إدارة تنمية مادة العلوم
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Weekly assessment

Question (1):- Use the following choices for the following questions :-

- 1- Hydrogen bonding. 2- Ionic. 3- Polar covalent.
4- Pure covalent. 5- Metallic. 6-Coordinate covalent.

(1) When the electro negativity difference between two atoms is 2, what type Of bond can be predicted?

(2) If two atoms are bonded as one of them donates a pair of electrons to the other, what type of bond is it?

(3) Which of the above bonds explain water's abnormally high boiling point?

(4) Which of the above bonds has its strength dependent on the number of valence electrons?

(5) A bond in which the bonding electrons spend the same time in the valence shell of each of the two bonded atoms?

(6) What is the type of bond in rod of iron?

Question two: Predict which solid in each of the following will have the higher melting point .Explain your answer.

(a) NaCl or CsCl

(b) $_{19}\text{K}$ or $_{20}\text{Ca}$

(c) MgCl_2 or AlCl_3

Question three: Which of the following compounds could be able to form hydrogen bonds between its molecules?

(1) CH_3OCH_3

(2) H_2S

(3) $\text{CH}_3\text{-NH}_2$

(4) NH_2OH

(5) PH_3

(6) CH_3COOH





الإدارة

إدارة تنمية مادة العلوم

Question four: -Give reasons for each of the following:

1- Boiling point of water is more than that of ammonia.

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2-The strength of metallic bond increases as the number of valence electrons increases.

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3-Sodium ($_{11}\text{Na}$) is softer than magnesium ($_{12}\text{Mg}$).

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