

برعاية معالي وزير التربية والتعليم

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وتوجيهات مساعد الوزير لشئون تطوير المناهج التعليمية
والمشرف على الادارة المركزية لتطوير المناهج

د/ اكرم حسن

اداءات وتقييمات

الصف الثاني الاعدادي

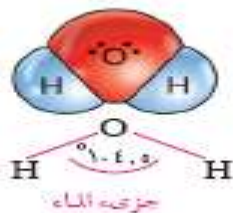
لجنة الاعداد والمراجعة

خبراء مكتب تنمية مادة العلوم

اشراف علمي

مستشار العلوم

د/ عزيزه رجب خليفة



الاختبارات الأسبوعية الخامس و السادس

الصف الثاني الإعدادي

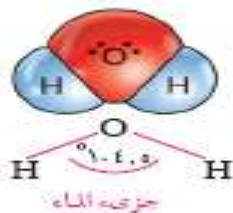
الدرس الثالث

Put (✓) or (X):

- 1- Alkali elements belong to class P. in the periodic table.
- 2- During the chemical reaction ,alkali elements lose an electron of its outer energy level turning into a negative ion.
- 3- Alkali are kept under water surface to prevent its reaction with air oxygen
- 4- The chemical activity increases by going down in group A1
- 5- Alkali reacts with water forming basic solutions
- 6- Liquefied nitrogen is used in preservation of the cornea of eyes.
- 7- Chlorine , Bromine and Astatine are found in nature.
- 8- Chlorine belongs to liquid halogens.
- 9- Bromine replaces Chlorine in its solutions.
- 10- Halogens elements belong to class P.

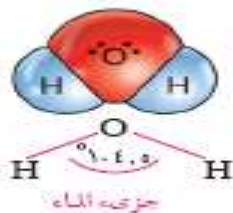
What happens when ..?

- 1- Putting a piece of Sodium in a glass of water.
- 2- Putting a piece of Sodium in kerosene
- 3- Adding drops of violet litmus solution to the formed solution from reacting sodium with water.
- 4- Passing chlorine gas into Sodium iodide
- 5- Putting a piece of Potassium in a tube with liquid bromine
- 6- Adding Bromine to Sodium Chloride solution.



Choose the correct answer:

- 1- Alkali metals distinguish with
(bad heat conductor - good electric conductor - high density - forming basic solution by reacting with water)
- 2- A metal alkali lies in the 3rd period ,its atomic number is
(3 - 5 - 11 - 19)
- 3- Group (1A) elements are called
(Alkali - Earth alkali - Halogens - Inert gas)
- 4- All the following elements are kept under kerosene EXEPT
(Sodium - Potassium - Rubidium - Lithium)
- 5- Sodium is more active than
(Cesium - Potassium - Rubidium - Lithium)
- 6- The less density alkali metals is
(Cs - K - Na - Li)
- 7- Alkali elements lie in the of the modern periodic table
(right - left - middle - beneath)
- 8- The alkali metal element with the largest atomic volume is
(Sodium - Potassium - Cesium - Lithium)
- 9- Reacting Sodium with water stronger the reacting water with
(Cesium - Potassium - Rubidium - Lithium)
- 10- The outer energy level of the alkali elements has Electron
(1 - 2 - 3 - 4)



11- All the following are properties of Alkali metals EXCEPT
 (good heat conductor - good electric conductor - hard to bend - forming positive ions during the reactions)

12- All of the following floats EXCEPT

(Rb - Na - K - Li)

13- Potassium is forming ion during the chemical reaction.

(K^- - K^+ - K^{++} - K^-)

14- is the highest alkali element in density.

(Cs - K - Na - Li)

15- The modern naming of the alkali elements is

(1 - 2 - 17 - 18)

16- The alkali elements has electron(s) in its outer shell .

(1 - 2 - 3 - 4)

17- When an element in group 1A reacts with another element in group 7A ,
 the If forming .

(acid - base - oxide - salt)

18- is used in preservation of the cornea of eyes.

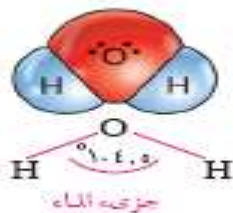
(Liquefied nitrogen - Liquefied sodium - Cobalt 60 - Silicon)

19- Halogens molecules consist ofatom(s)

(one - two - three - four)

20- Bromine replaces In its solutions .

(chlorine - fluorine - iodine - Argon)



Explain with chemical equation of the reaction of

- 1- Potassium with water
- 2- Sodium with water
- 3- Sodium with chlorine
- 4- Chlorine with Potassium bromide
- 5- bromine with Potassium iodide.

